

Objective
Paper Code**6481**

Intermediate Part First (New Scheme)

CHEMISTRY (Objective) GROUP - I

Time: 20 Minutes

Marks: 17

Q.No.1 You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill the relevant circle in front of that question number on computerized answer sheet. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero marks in that question. Attempt as many questions as given in objective type question paper and leave other circles blank.

S.#	Questions	A	B	C	D
1	The largest number of molecules are present in:	3.6g of H ₂ O	4.8g of C ₂ H ₅ OH	2.8g of CO	5.4g of N ₂ O ₅
2	The number of moles of CO ₂ which contain 8.0g of oxygen:	0.25	0.50	1.00	1.50
3	The comparative rates at which the solutes move in paper chromatography, depend on:	The size of paper	R _f values of solutes	Temperature of the experiment	Size of chromatographic tank used
4	Pressure remaining constant, at which temperature the volume of a gas will become twice of what it is at 0°C?	546°C	200°C	546K	273K
5	The molar volume of CO ₂ is maximum at:	STP	127°C and 1 atm	0°C and 2 atm	273°C and 2 atm
6	Acetone and chloroform are soluble in each other due to:	Intermolecular hydrogen bonding	Instantaneous dipole	Ion-dipole interaction	All of these
7	The molecules of CO ₂ in dry ice form the:	Ionic crystals	Covalent crystals	Molecular crystals	Any type of crystal
8	Orbitals having same energy are called:	Hybrid orbitals	Valence orbitals	Degenerate orbitals	d-orbitals
9	When 6d orbital is complete, the entering electron goes into:		7s	7p	7d
10	The hydrogen halides that has the highest percentage of ionic character:	HCl	HBr	HF	HI
11	The number of bonds in nitrogen molecule is:	One sigma and one pi	Three sigma only	One sigma and two pi	Two sigma and one pi
12	Calorie is equivalent to:	0.4184J	41.84J	4.184J	418.4J
13	The pH of 10 ⁻³ mol dm ⁻³ of an aqueous solution of H ₂ SO ₄ is:	3.0	2.7	2.0	1.5
14	The molal boiling point constant is the ratio of the elevation in boiling point is:	Molarity	Molality	Mole fraction of solute	Mole fraction of solvent
15	Molarity of pure water is:	1	18	55.5	6
16	If the salt bridge is not used between two half cells, then the voltage:	Drops to zero	Decreases rapidly	Decreases slowly	Does not change
17	The unit of the rate constant is the same as that of the rate of reaction is:	Zero order reaction	First order reaction	Second order reaction	Third order reaction

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