

Roll No. : _____

Objective
Paper Code**6484**

Intermediate Part First

CHEMISTRY (Objective) GROUP - II

Time: 20 Minutes

Marks: 17



Q.No.1

You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill the relevant circle in front of that question number on computerized answer sheet. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero marks in that question. Attempt as many questions as given in objective type question paper and leave other circles blank.

S.#	Questions	A	B	C	D
1	Indicate the catalyst used for the reaction: $\text{HCOOH} \rightarrow \text{H}_2\text{O} + \text{CO}$	Cu	<u>MnO₂</u>	Pt	<u>Al₂O₃</u>
2	If the salt bridge is not used between two half cells, then the voltage:	Decreases rapidly	<u>Decreases slowly</u>	Does not change	Drops to zero
3	The reaction at cathode during the electrolysis of dil. H ₂ SO ₄ with Pt electrodes is:	Oxidation	<u>Reduction</u>	Both oxidation and reduction	Neither oxidation nor reduction
4	Which solution has the highest boiling point?	5.85% solution of NaCl	<u>18.0% solution of glucose</u>	6.0% solution of urea	All have the same boiling point
5	When H ₂ S is added to HCl aqueous solution, the ionization of H ₂ S:	Increases	Remains constant	Decreases	Increases rapidly
6	An excess of aqueous silver nitrate is added to aqueous barium chloride and precipitate is removed by filtration. What are the main ions in the filtrate?	Ag ⁺ and NO ₃ ⁻ only	<u>Ba²⁺ and NO₃⁻ only</u>	Ba ²⁺ , NO ₃ ⁻ and Cl ⁻	Ag ⁺ , Ba ²⁺ and NO ₃ ⁻
7	The change in heat energy of a chemical reaction at constant temperature and pressure is called:	Internal energy change	<u>Bond energy</u>	Enthalpy change	Heat of sublimation
8	If an endothermic reaction is allowed to take place very rapidly in the air, the temperature of the surrounding air:	Decreases	<u>Increases</u>	Remains constant	Remains unchanged
9	The number of bonds in nitrogen molecule is:	One sigma and one pi	<u>One sigma and two pi</u>	Three sigma only	Two sigma and one pi
10	In the ground state of an atom, the electron is present:	In the nucleus	<u>In the second shell</u>	Nearest to the nucleus	Farthest from the nucleus
11	NaF and MgO are isomorphs of each other and exist in:	Tetragonal form	<u>Rhombohedral form</u>	Orthorhombic form	Cubic form
12	London dispersion forces are the only forces present among the:	Molecules of water in liquid state	<u>Atoms of He in gaseous state at high temperature</u>	Molecules of hydrogen chloride gas	Molecules of solid iodine
13	The value of R in NmK ⁻¹ mol ⁻¹ is:	1.987	<u>8.3143</u>	0.0821	62.4
14	A real gas obeying van der Waals equation will resemble ideal gas if:	Both "a" and "b" are small	<u>Both "a" and "b" are large</u>	"a" is small and "b" is large	"a" is large and "b" is small
15	The stationary phase in adsorption chromatography is:	Solid	<u>Water</u>	Organic liquid	Gas
16	The mass of water formed when 2g of H ₂ and 64g of O ₂ are combined together is:	68g	<u>36g</u>	18g	66g
17	27g of Al will react completely with how much mass of O ₂ to produce Al ₂ O ₃ ?	32g of oxygen	<u>24g of oxygen</u>	16g of oxygen	8g of oxygen

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