Objective Paper Code

8488

Intermediate Part Second

CHEMISTRY (Objective) GROUP - II

Time: 20 Minutes

Marks: 17



You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill the relevant circle in front of that question number on computerized answer sheet. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero marks in that question. Attempt as many questions as given in objective type question person and leave of the circles. Q.No.1 objective type question paper and leave other circles blank.

S.#	Questions	A	В	С	D
1	The electrophile in aromatic sulphonation is:	H ₂ SO ₄	HSO ₄	SO ₃	SO ₃ ⁺
2	When cyanogen chloride $(C\ell - CN)$ is made to react with ethyl magnesium bromide the product formed is:	CH ₃ – CN	CH ₃ – CH ₂ – CN	CH ₃ – CH ₂ – CH ₂ – CN	CH ₂ = CH – CN
3	Which compound will have the maximum repulsion with water?	C ₆ H ₆	C ₂ H ₅ OH	CH ₃ CH ₂ CH ₂ – OH	CH ₃ = O - CH ₃
4	It will have the highest boiling point:	Mathanal	Ethanal	Propanal	2-Hexanone
5	Which reagent is used to reduce a carboxylic group to an alcohol?	H ₂ / Ni	H ₂ / Pt	NaBH ₄	LiAℓH ₄
6	Nitrogeneous bases is not present in RNA:	Cytosine	Adenine	Thiamine	Uracil
7	Phosphorus helps the growth of:	Root	Leave	Stem	Seed
8	Methane has a mean residence time in atmosphere about:	3 – 7 years	4 – 7 years	5 – 7 years	6 – 7 years
9	Newspaper can be recycled again and again by how many times?	2	3	4	5
10	Select the two normal elements are present in seventh period:	Rb, Sr	Cs, Ba	Fr, Ra	La , Hf
11	It does not belong to alkaline earth metals:	Be	Ra	Ba	Rn
12	The chief ore of aluminum is:	Na ₃ AℓF ₆	$A\ell_2O_3 \cdot 2H_2O$	Aℓ ₂ O ₃	$A\ell_2O_3 \cdot H_2O$
13	Birkeland-Eyde process used for the preparation of:	HNO ₃	H ₂ SO ₄	C ₆ H ₆	НСНО
14	The anhydride of HCℓO ₄ is:	${\rm C}\ell{\rm O}_3$	CℓO ₂	Cℓ ₂ O ₅	Cℓ ₂ O ₇
15	It is a typical transition metal:	Sc	Y	Co	Ra
16	Linear shape is associated with which set of hybrid orbitals?	sp	sp ²	sp ³	dsp ²
17	Synthetic rubber is made by polymerization of:	Chloroform	Acetylene	Divinyl acetylene	Chloroprene